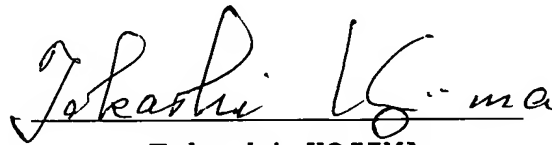


C E R T I F I C A T I O N

I, Takashi KOJIMA of Ginza Ohtsuka Bldg., 2F, 16-12, Ginza 2-chome, Chuo-ku, Tokyo, Japan, hereby certify that I am the translator of the accompanying certified official copy of the documents in respect of an application for a patent filed in Japan on the 28th of February, 2001 and of the official certificate attached thereto, and certify that the following is a true and correct translation to the best of my knowledge and belief.

Dated this 13th day of January, 2004

  
Takashi KOJIMA

(Translation)

PATENT OFFICE  
JAPANESE GOVERNMENT

This is to certify that the annexed is a true copy of the following application as filed with this Office.

Date of Application : February 28, 2001

Application Number : Japanese Patent Application  
No. 2001-053903

Applicant(s) : Shin-Etsu Chemical Co., Ltd.

August 24, 2001

Commissioner,  
Patent Office

Kozo OIKAWA  
(sealed)

Certificate No. 2001-3076020

2001-053903

[Paper]

APPLICATION FOR PATENT

[Reference Number]

13012

[Application Date]

February 28, 2001

[Destination]

The Commissioner of the Patent Office  
Kozo OIKAWA

[International Patent Classification]

C01B 33/12

[Inventor]

[Address]

c/o Gunma Complex,  
Shin-Etsu Chemical Co., Ltd.,  
2-13-1, Isobe, Annaka-shi, Gunma-ken,  
Japan

[Name]

Hirofumi FUKUOKA

[Inventor]

[Address]

c/o Gunma Complex,  
Shin-Etsu Chemical Co., Ltd.,  
2-13-1, Isobe, Annaka-shi, Gunma-ken,  
Japan

[Name]

Satoru MIYAWAKI

[Inventor]

[Address]

c/o Isobe Service Inc.,  
2-13-1, Isobe, Annaka-shi, Gunma-ken,  
Japan

[Name]

Kenji Oōka

[Inventor]

[Address]

c/o Gunma Complex,  
Shin-Etsu Chemical Co., Ltd.,  
2-13-1, Isobe, Annaka-shi, Gunma-ken,  
Japan

[Name]

Susumu UENO

[Inventor]

[Address]

c/o Gunma Complex,  
Shin-Etsu Chemical Co., Ltd.,  
2-13-1, Isobe, Annaka-shi, Gunma-ken,  
Japan

[Name]

Mikio ARAMATA

[Inventor]  
[Address] c/o Shin-Etsu Chemical Co., Ltd.,  
6-1, Otemachi 2-chome, Chiyoda-ku,  
Tokyo, Japan  
[Name] Takeshi FUKUDA

[Applicant]  
[Identification Number] 000002060  
[Name] Shin-Etsu Chemical Co. Ltd.

[Agent]  
[Identification Number] 100079304  
[Patent Attorney]  
[Name] Takashi KOJIMA

[Agent]  
[Identification Number] 100103595  
[Patent Attorney]  
[Name] Yuko NISHIKAWA

[Official Fee]  
[Deposit Account Number] 003207  
[Amount of Fee] 21000

[List of Documents Attached]  
[Document] Specification 1  
[Document] Drawing 1  
[Document] Abstract 1  
[Necessity of Proof] Yes

2001-053903

[SPECIFICATION]

[TITLE OF THE INVENTION] Silicon Oxide Powder and Making  
Method

5 [CLAIMS]

[Claim 1] A silicon oxide powder characterized by being represented by the general formula:  $\text{SiO}_x$  wherein  $x$  is in the range of  $1.05 \leq x \leq 1.5$  and having a BET specific surface area of 5 to 300  $\text{m}^2/\text{g}$ .

10 [Claim 2] A method for preparing the silicon oxide powder of claim 1, characterized by heating a raw material powder mixture containing at least a silicon dioxide powder in an inert gas atmosphere or in vacuum at a temperature of 1,100 to 1,600°C to generate  $\text{SiO}$  gas, continuously or  
15 intermittently feeding oxygen gas to the  $\text{SiO}$  gas to form a gas mixture, and depositing the gas mixture on a surface of a cooled substrate.

[Claim 3] The method of claim 2 characterized in that the temperature of the cooled substrate surface is from 200 to  
20 400°C.

[DETAILED DESCRIPTION OF THE INVENTION]

[0001]

[Technical Field of the Invention]

This invention relates to a silicon oxide powder for  
25 use as a negative electrode active material in lithium ion secondary cells, and a method for preparing the same.

[0002]

[Prior Art]

Silicon oxide represented by  $\text{SiO}$  is a well-know  
30 material. By taking advantage of its chemical activity, it is used, for example, in the synthesis of industrially useful alkylhalosilanes (Gary N. Bokerman et al., USP 5,051,247) and the direct synthesis of siloxanes (Peter L. Timms and William N. Rowlands, EPA 0406000A2 "Polysiloxane oils and process for  
35 their preparation"). It is also reported that magnesium

silicide is prepared by reacting silicon oxide with magnesium at low temperatures (Fuglein, E; and Schubert, U, Chem. Mater. 1999, 11, 865-866).

Meanwhile, it is reported that the use of  $\text{SiO}_x$  as a negative electrode active material in lithium ion secondary cells facilitates occlusion and release of lithium ions (JP-A 9-7638). Silicon oxide is thus expected to find an extra application as the negative electrode active material in lithium ion secondary batteries.

10 [0003]

[Problem to be solved by the Invention]

The use of silicon oxide  $\text{SiO}_x$  powder as the negative electrode material in lithium ion secondary cells leads to a drastically increased initial capacity, but poor cycle performance.

15

[0004]

The present invention has been accomplished under the above circumstances. An object of the invention is to provide a silicon oxide powder which is used as the negative electrode material to construct a lithium ion secondary cell capable of maintaining a high capacity without detracting from cycle performance, and a method for preparing the silicon oxide powder.

20

[0005]

25 [Means for solving the problem and Embodiment of the Invention]

The inventors have earnestly studied to attain the above objects and have found that when a silicon oxide powder of the general formula:  $\text{SiO}_x$  wherein  $x$  has a value within a certain range and having a certain specific surface area is used as the negative electrode material in a lithium ion secondary cell, the cell has a high capacity and excellent cycle performance. The conditions under which a silicon oxide powder of a specific composition having physical properties in the desired range can be prepared have been empirically established.

30

35

[0006]

Accordingly, the present invention provides:

(1) a silicon oxide powder characterized by being represented by the general formula:  $\text{SiO}_x$  wherein  $x$  is in the range of

5  $1.05 \leq x \leq 1.5$  and having a BET specific surface area of 5 to 300  $\text{m}^2/\text{g}$ , and

(2) a method for preparing the above silicon oxide powder, characterized by heating a raw material powder mixture containing at least a silicon dioxide powder in an inert gas  
10 atmosphere or in vacuum at a temperature of 1,100 to 1,600°C to generate  $\text{SiO}$  gas, continuously or intermittently feeding oxygen gas to the  $\text{SiO}$  gas to form a gas mixture, and depositing the gas mixture on a surface of a cooled substrate.

15 [0007]

Following is a detailed description of the invention.

The silicon oxide powder of the invention is characterized by being represented by the general formula:  $\text{SiO}_x$  wherein  $x$  is in the range of  $1.05 \leq x \leq 1.5$  and having a  
20 specific surface area of 5 to 300  $\text{m}^2/\text{g}$  as measured by the BET adsorption method.

[0008]

When a silicon oxide powder of the general formula:  $\text{SiO}_x$  wherein  $x$  is less than 1.05 is used as the negative  
25 electrode material in a lithium ion secondary cell, the initial capacity of the cell is increased at the sacrifice of cycle performance. If  $x$  is more than 1.5, the cycle performance is improved at the sacrifice of charge-discharge capacity. Preferably  $x$  is  $1.1 \leq x \leq 1.3$ .

30 [0009]

It is not well understood why the performance of a lithium ion secondary cell changes with the value of  $x$  in the formula:  $\text{SiO}_x$  representative of the silicon oxide powder. It is presumed that as the  $x$  value that is the quantity of  
35 atomically dispersed oxygen increases, the silicon oxide is reduced in percent volume expansion upon occlusion of lithium

ions. This restrains the disintegration and powdering of the negative electrode material by repetition of charge/discharge cycles. As a result, cycle performance is improved.

[0010]

5           A BET surface area of less than 5 m<sup>2</sup>/g indicates that silicon oxide particles have a lower surface activity. When this silicon oxide powder is used as the negative electrode material in a lithium ion secondary cell, the amount of lithium ions migrating into and out of the silicon oxide  
10 powder is reduced and as a result, the charge/discharge capacity is reduced. If the BET surface area is more than 300 m<sup>2</sup>/g, silicon oxide particles are so vulnerable to surface oxidation that the proportion of SiO<sub>2</sub> increases and hence, the value of x increases, resulting in a lowering of  
15 charge/discharge capacity. The preferred range of BET specific surface area is 10 to 200 m<sup>2</sup>/g.

[0011]

          According to the invention, the silicon oxide powder defined above is prepared by heating a raw material powder  
20 mixture containing at least a silicon dioxide powder in an inert gas atmosphere or in vacuum at a temperature in the range of 1,100 to 1,600°C to generate SiO gas, continuously or intermittently feeding oxygen gas to the SiO gas to form a gas mixture, and depositing the gas mixture on a surface of a  
25 cooled substrate.

[0012]

          The raw material powder mixture containing a silicon dioxide powder is typically a mixture of silicon dioxide and a powder capable of reducing the silicon dioxide.  
30 Illustrative of the reducing powder are metallic silicon and carbon-containing powders. Of these, metallic silicon is preferred for increased reactivity and yields.

          The metallic silicon used herein is not critical although metallic silicon species of high purity such as  
35 semiconductor grade Si, ceramic grade Si and chemical grade Si are preferred because of the increased purity of the resultant silicon oxide powder.



The mixing proportion of silicon dioxide and a reducing powder is determined as appropriate as long as complete reduction is ensured.

[0013]

5 In the present invention, the mixture of silicon dioxide and a reducing powder is heated and maintained at a temperature in the range of 1,100 to 1,600°C, and preferably 1,200 to 1,500°C.

10 If the raw material powder mixture is heated below 1,100°C, reaction proceeds with difficulty, resulting in a lower productivity. At a heating temperature above 1,600°C, the mixture melts and loses reactivity and a choice of the furnace material becomes a concern.

15 The atmosphere within the furnace may be either an inert gas or vacuum although the reaction in vacuum is advantageous in that reactivity becomes higher on account of thermodynamics so that reaction can proceed at lower temperatures.

[0014]

20 The SiO gas generated by heating the raw material powder is then fed through a feed conduit to a deposition chamber for depositing a silicon oxide powder. The feed conduit is preferably heated and maintained at a temperature of 1,000 to 1,300°C, more preferably 1,100 to 1,200°C. If  
25 the feed conduit is below 1,000°C, the SiO gas may deposit and build up on the inner wall of the conduit, which becomes a disturbance to operation, eventually retarding stable continuous operation. Heating the conduit above 1,300°C achieves no additional effects and can invite an increase of  
30 power consumption.

[0015]

While the SiO gas is carried to the deposition chamber, oxygen gas is fed thereto. The oxygen gas may be fed as such or an inert gas containing oxygen gas be fed. By  
35 controlling the flow rate and feed time of oxygen gas, the x value in the resulting silicon oxide powder (SiO<sub>x</sub>) can be adjusted.

It is not critical how to feed the oxygen gas, and the oxygen gas may be fed continuously or intermittently or otherwise. Any appropriate oxygen gas feed method may be selected in accordance with a particular purpose. When  
5 oxygen gas is fed to and mixed with SiO gas, the temperature is preferably 800 to 1,200°C, and especially 900 to 1,100°C.

[0016]

A gas mixture obtained by feeding oxygen gas to SiO gas is contacted with a surface of a substrate which is  
10 disposed within the deposition chamber and cooled with a coolant whereby a silicon oxide powder (SiO<sub>x</sub>) having a predetermined value of x deposits on the substrate.

The temperature of the substrate at the surface is, though not critical, preferably 200 to 400°C. At a substrate  
15 surface temperature below 200°C, the resulting silicon oxide powder may have a BET specific surface area in excess of 300 m<sup>2</sup>/g and a lower purity. On the other hand, at a substrate surface temperature above 400°C, the resulting silicon oxide powder may have a BET specific surface area of less than 5  
20 m<sup>2</sup>/g and a lower activity.

[0017]

The coolant for cooling the substrate is, though not critical, selected from liquids such as water and heat transfer media and gases such as air and nitrogen gas. The  
25 type of the substrate is not critical although the use of high-melting metals such as stainless steel, molybdenum and tungsten is preferred for ease of working.

[0018]

It is not well understood why the BET specific surface  
30 area of silicon oxide powder changes with the substrate surface temperature. As the substrate surface temperature becomes higher, deposits become sufficiently active on their surface to fuse them together for consolidation, leading to a lowering of BET specific surface area.

35 [0019]

The silicon oxide powder thus deposited on the substrate is recovered by a suitable means such as a scraper.

After recovery, the silicon oxide powder is pulverized to a desired particle size in a suitable means such as a ball mill, if necessary.

[0020]

5           Now an apparatus for use in the practice of the inventive method for producing a silicon oxide powder is described. Referring to FIG. 1, there is illustrated an apparatus according to one embodiment of the invention, generally designated at 1.

10           The apparatus includes a reaction furnace 10 and a deposition chamber 20. The reaction furnace 10 and the deposition chamber 20 are interconnected by a transfer conduit 30 having a heater 31 built therein.

          Disposed in the reaction furnace 10 is a muffle 11.  
15   The muffle 11 is surrounded by a heater 12. A heat insulator 13 is disposed around the heater 12.

[0021]

          The muffle 11 defines a reaction chamber 14 therein. Disposed within the reaction chamber 14 is a container 16 for  
20   receiving a raw material powder mixture 15 containing a silicon dioxide powder. The heater 12 is operable to heat the raw material powder mixture 15 in the container 16 at a predetermined temperature to thereby generate a silicon oxide gas which is then fed into the deposition chamber 20 through  
25   the transfer conduit 30 connected to the muffle 11.

[0022]

          Disposed in the deposition chamber 20 are a substrate 21 and a heater 22 for heating the interior of the chamber 20. Connected to the deposition chamber 20 are a gas inlet  
30   tube 23 for feeding oxygen gas into it and a vacuum pump 24 for evacuating and maintaining in vacuum the interior of the apparatus.

          The substrate 21 is provided in its interior with coolant channels (not shown) to which a coolant inlet tube 25  
35   and a coolant outlet tube 26 are connected.

          The silicon oxide gas which has been introduced into the deposition chamber 20 through the feed conduit 30 is

brought in contact with the surface of the cooled substrate 21 whereby it is cooled and deposited as silicon oxide particles.

[0023]

5 [EXAMPLE]

Examples of the invention and Comparative Examples are given below by way of illustration and not by way of limitation.

[0024]

10 [Example 1]

Using the apparatus 1 shown in FIG. 1, a silicon oxide  $\text{SiO}_x$  powder was prepared. The raw material powder mixture 15 used was an equimolar mixture of a silicon dioxide powder (BET surface area  $200 \text{ m}^2/\text{g}$ ) and a metallic silicon powder of 15 the ceramic grade (BET surface area  $4 \text{ m}^2/\text{g}$ ).

[0025]

First, 200 g of the raw material powder mixture 15 was placed in the reaction furnace 10 whose muffle 11 had a volume of  $6,000 \text{ cm}^3$ . The vacuum pump 24 was operated to 20 evacuate the reaction furnace 10 to a vacuum of 0.1 Torr or lower. The heater 12 was actuated to heat and maintain the container at a temperature of  $1,350^\circ\text{C}$ . With the heater 31 on, the transfer conduit 30 was heated and maintained at a temperature of  $1,100^\circ\text{C}$ .

25 The heater 22 in the deposition chamber 20 was then actuated to heat the chamber 20 at a temperature of  $900^\circ\text{C}$ . At the same time, water was flowed at a rate of  $5.0 \text{ NL/min}$  through the channels in the stainless steel substrate 21 (surface area of  $200 \text{ cm}^2$ ).

30 [0026]

An argon gas containing 20% of oxygen gas was continuously fed at a rate of  $50 \text{ cc/min}$  from the gas inlet tube 23 into the deposition chamber 20.

Under these conditions, the surface of the substrate 35 21 was maintained at a temperature of about  $280^\circ\text{C}$ .

Operation was continued for 5 hours, after which 160 g of SiO<sub>x</sub> in a black mass form deposited on the surface of the substrate 21. The mass deposit was recovered and pulverized for 5 hours in a ball mill, obtaining a silicon oxide powder.

5 This silicon oxide powder was an amorphous powder having a BET specific surface area of 210 m<sup>2</sup>/g and represented by the general formula: SiO<sub>x</sub> (wherein x=1.22).

[0027]

10 Next, 100 parts by weight of the silicon oxide powder thus obtained, 90 parts by weight of graphite as a conductive agent, and 20 parts by weight of polyvinylidene fluoride as a binder were kneaded together (in N-methylpyrrolidone solvent). A portion of this compound was applied to a stainless steel mesh, bonded thereto under pressure, and  
15 dried at 120°C overnight in a vacuum dryer, obtaining an electrode containing silicon oxide. (The preparation conditions and the physical properties of silicon oxide are shown in Table 1.)

[0028]

20 The charge/discharge performance of the electrode was examined. A lithium ion secondary cell for the rating of the electrode was constructed using a lithium foil as the counter electrode. The nonaqueous electrolyte used was a nonaqueous electrolyte solution of lithium phosphorus hexafluoride in a  
25 1/1 mixture of ethylene carbonate and 1,2-dimethoxyethane in a concentration of 1 mol/liter. The separator used was a microporous polyethylene film of 30 μm thick.

[0029]

30 The lithium ion secondary cell thus constructed was allowed to stand overnight at room temperature. Using a secondary cell charge/discharge tester (Nagano K.K.), a charge/discharge test was carried out on the cell under conditions including a constant current density of 0.5 mA/cm<sup>2</sup>, a discharge termination voltage of 0.003 V and a  
35 charge termination voltage 1.800 V (10 cycles). The results are shown in Table 2.

[0030]

[Examples 2 to 4]

Silicon oxide powders represented by the general formula:  $\text{SiO}_x$  were prepared as in Example 1 except that the temperature within the deposition chamber 20, the flow rate of water through the substrate 21, and the flow rate of 20% oxygen-containing argon gas were changed as shown in Table 1.

Using the silicon oxide powders thus obtained, a charge/discharge test was carried out as in Example 1. The results of the charge/discharge test are shown in Table 2.

[0031]

[Comparative Examples 1 to 4]

Silicon oxide powders represented by the general formula:  $\text{SiO}_x$  were prepared as in Example 1 except that the temperature within the deposition chamber 20, the flow rate of water through the substrate 21, and the flow rate of 20% oxygen-containing argon gas were changed as shown in Table 1.

Using the silicon oxide powders thus obtained, a charge/discharge test was carried out as in Example 1. The results of the charge/discharge test are shown in Table 2.

[0032]

[Table 1]

	SiO <sub>x</sub> preparing conditions				SiO <sub>x</sub>	
	Deposition chamber temp. (°C)	Water flow rate (NL/min)	Ar+20%O <sub>2</sub> gas flow rate (cc/min)	Substrate surface temp. (°C)	X value	BET surface area (m <sup>2</sup> /g)
Example 1	900	5.0	50	280	1.22	210
Example 2	900	5.0	200	280	1.45	200
Example 3	800	7.0	50	210	1.32	280
Example 4	1000	5.0	50	360	1.15	35
Comparative Example 1	900	5.0	0	280	1.01	210
Comparative Example 2	900	5.0	300	280	1.65	230
Comparative Example 3	700	7.0	50	160	1.30	350
Comparative Example 4	1050	5.0	50	420	1.10	3

[0033]

[Table 2]

	Initial charge/discharge			10th charge/discharge		
	Charge capacity (mAh/g)	Discharge capacity (mAh/g)	Capacity loss (%)	Charge capacity (mAh/g)	Discharge capacity (mAh/g)	Discharge retentivity (%)
Example 1	1150	742	35	570	566	76
Example 2	1010	720	29	560	557	77
Example 3	1045	810	22	577	570	70
Example 4	1180	780	34	590	583	75
Comparative Example 1	1320	610	53.8	235	222	36
Comparative Example 2	400	315	21	260	255	81
Comparative Example 3	312	240	23	190	185	77
Comparative Example 4	420	310	26	240	235	76

[0034]

[Effect of the Invention]

There has been described a silicon oxide powder having the general formula:  $\text{SiO}_x$  wherein  $x$  is controlled in a specific range, which is used as a negative electrode material to construct a lithium ion secondary cell having a high capacity and improved cycle performance. The silicon oxide

[BRIEF DESCRIPTION OF THE DRAWINGS]

[FIG. 1]

A schematic cross-sectional view of an apparatus for preparing silicon oxide according to one embodiment of the invention.

[Explanation of Symbols]

- 1 apparatus for producing a silicon oxide powder
- 10 reaction furnace
- 11 muffle
- 12 heater
- 13 heat insulator
- 14 reaction chamber
- 15 raw material powder mixture
- 16 raw material container

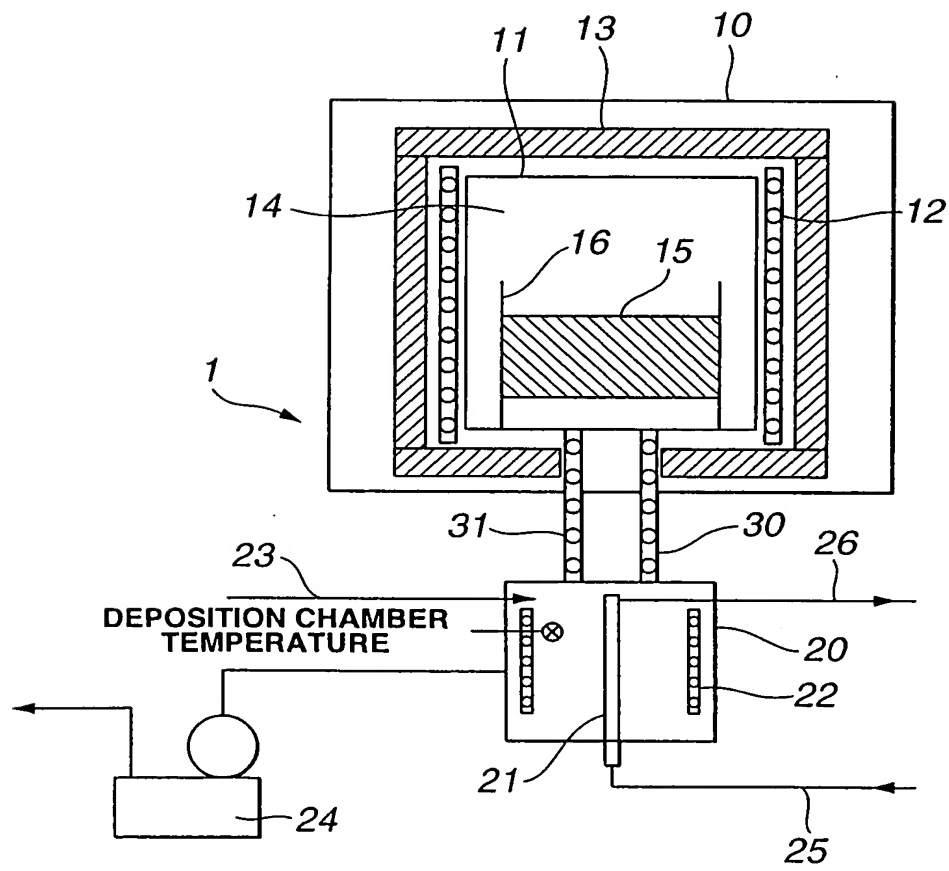
20 deposition chamber  
21 substrate  
22 heater  
23 gas inlet tube  
5 24 vacuum pump  
25 coolant inlet tube  
26 coolant outlet tube  
30 transfer conduit  
31 heater



2001-053903

[DRAWING]

[FIG. 1]



2001-053903

[ABSTRACT]

[Object]

To provide a silicon oxide powder which is used as the negative electrode material to construct a lithium ion secondary cell capable of maintaining a high capacity without detracting from cycle performance.

[Means for Solution]

A silicon oxide powder represented by the general formula:  $\text{SiO}_x$  wherein  $x$  is in the range of  $1.05 \leq x \leq 1.5$  and having a BET specific surface area of 5 to 300  $\text{m}^2/\text{g}$ .

[Selected Drawing]      None